

Name

Key

Per

LEWIS STRUCTURES AND VSEPR PRACTICE

Compound	Lewis Dot Structure	Geometry	Drawing	Polarity
				Y or N
F ₂ 14	$\begin{array}{c} \times \times \\ \times \times \end{array} \text{F} - \text{F} \begin{array}{c} \times \times \\ \times \times \end{array}$	linear	F-F	N
BI ₃ 24	$\begin{array}{c} \times \times \\ \times \times \end{array} \text{I} - \text{B} - \text{I} \begin{array}{c} \times \times \\ \times \times \end{array}$	t. planar		N
CCl ₄ 32	$\begin{array}{c} \times \times \\ \times \times \end{array} \text{Cl} \\ \times \times \\ \times \times \end{array} \text{C} - \text{C} \begin{array}{c} \times \times \\ \times \times \end{array} \text{Cl} \\ \times \times \\ \times \times$	tetra- hedral		N
H ₂ O 8	H - $\begin{array}{c} \times \times \\ \times \times \end{array} \text{O} - \text{H}$	Bent 105		Y
CO ₂ 16	$\begin{array}{c} \times \times \\ \times \times \end{array} \text{O} = \text{C} = \text{O} \begin{array}{c} \times \times \\ \times \times \end{array}$	linear	O=C=O	N
CO 10	$\times \text{C} \equiv \text{O} \times$	linear	C≡O	Y
NH ₃ 8	H - $\begin{array}{c} \times \times \\ \times \times \end{array} \text{N} - \text{H} \\ \times \times$	t. pyramidal		Y
SO ₄ ²⁻	$\left[\begin{array}{c} \times \times \\ \times \times \end{array} \text{O} \\ \times \times \\ \times \times \end{array} \text{S} - \text{O} \begin{array}{c} \times \times \\ \times \times \end{array} \right]^{-2}$	tetra- hedral		N

Compound	Lewis Dot Structure	Geometry	Drawing	Polarity
				Y or N
AsF ₃ 26		t. pyramidal		Y
SO ₃ 24		t. planar		N
SO ₂ 18		Bent 117		Y
C ₂ H ₂ (2) 10	H-C≡C-H	linear x2	H-C≡C-H	N
C ₂ H ₅ OH (3) 20		C1 tetra C2 tetra O Bent 105		Y
CH ₃ COOH (3) 24		C1 tetra C2 t. planar O Bent 105		Y
CH ₃ COCH ₃ (3) 24		C1 tetra C2 tri planar C3 tetra		Y