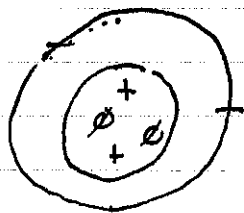


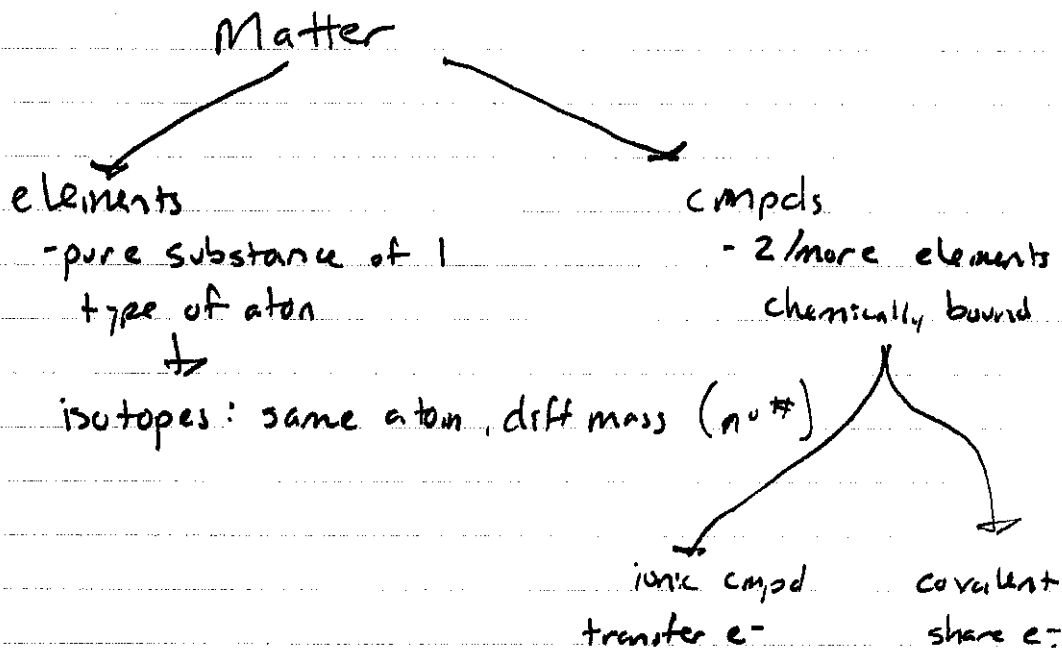
Macromolecules

Basic Chemistry

Atoms: basic unit of matter



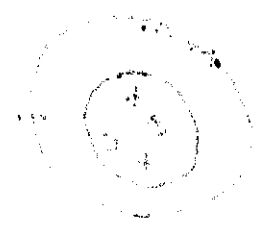
nucleus: protons (+)
neutrons (0)
cloud: electrons (-)



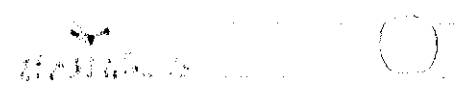
Handwritten title or section header.

Handwritten text line.

Handwritten text on the left side, possibly a list or notes.



Handwritten text below the circular diagram.



Handwritten text on the left side, below the first list.

Handwritten text in the middle, below the circular diagram.



Handwritten text below the second diagram.

Handwritten text on the left side, at the bottom.

Handwritten text in the middle, at the bottom.

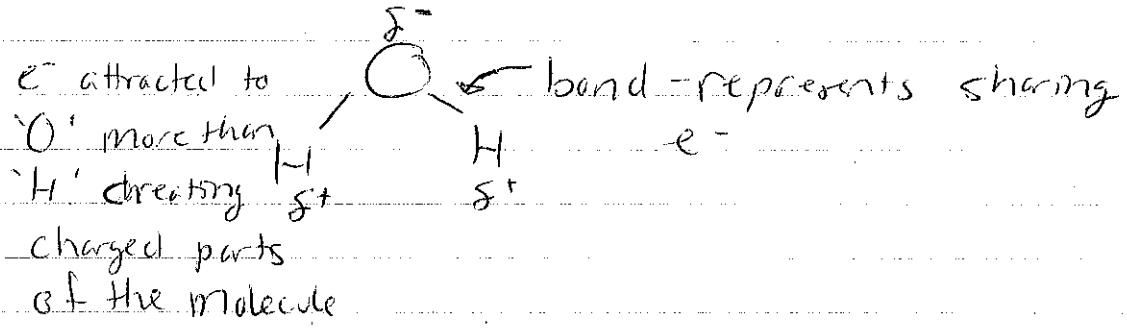
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Water : How does the structure of H₂O contribute to its unique properties?

① Polarity

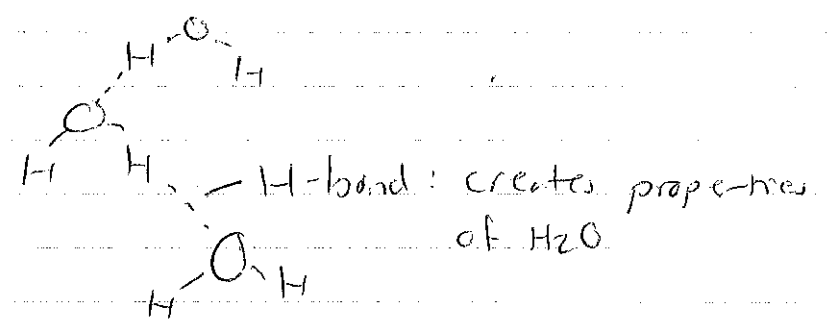
def: uneven distribution of charges (e⁻) around a molecule

creates
↓



② Hydrogen bonding

def: positively charged H on 1 H₂O is attracted to negatively charged 'O' on another H₂O



③ cohesion

def: attraction b/w molecules of same substance

explains: bead formation
surface tension

⑥ adhesion

def: attraction b/w molec. of
diff substances

-explains: meniscus

capillary action

transpiration

⑦ heat capacity

def: amount of heat energy
required to increase H_2O
temp is relatively high

explains: regulating human temp
bodies of water

Solutions: How does H_2O polarity influence
its properties as a solvent?

- mixtures: 2+ elements / compds. physically
mixed



- Solutions:

def: one substance is dissolved in another
(salt water)

evenly distributed

- solute: being dissolved

- solvent: dissolver

Buffers: method of controlling pH changes

importance: human body (blood)
oceans (global climate change)

Macromolecules Vocab

Atom

* Hydrogen bond / Polarity

Cohesion

Adhesion

Mixture

Solute

Solvent

* Solution

pH scale

Acid

Base