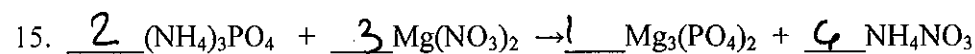
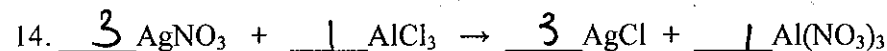
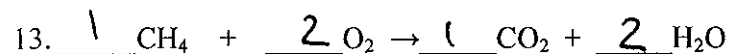
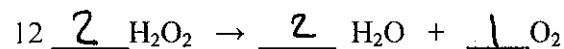
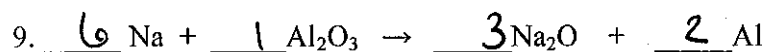
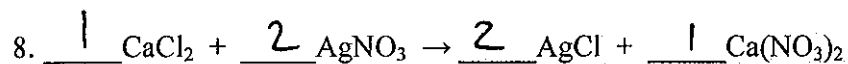
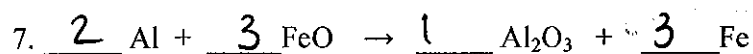
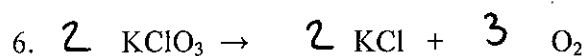
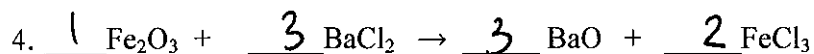
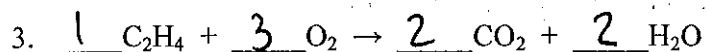
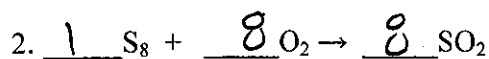
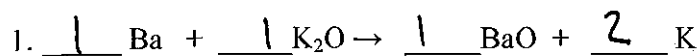
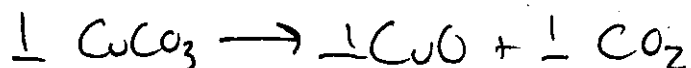


### More Balancing Practice

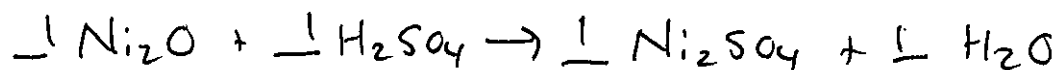


Write balanced equations for the following reactions.

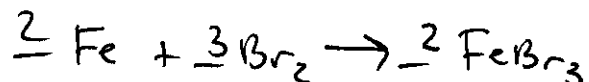
Copper (II) carbonate forming copper (II) oxide and carbon dioxide.



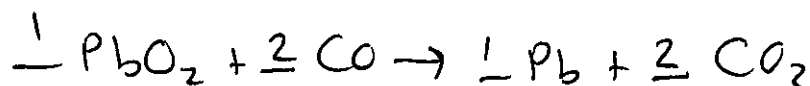
Nickel (I) oxide reacting with sulfuric acid ( $\text{H}_2\text{SO}_4$ ) to form nickel (I) sulfate and water.



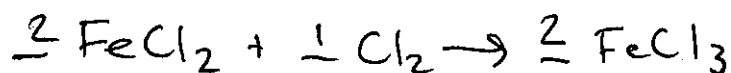
Iron and bromine reacting to give iron(III) bromide.



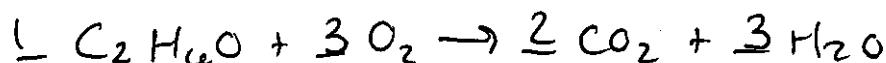
Lead (IV) oxide and carbon monoxide forming lead metal and carbon dioxide.



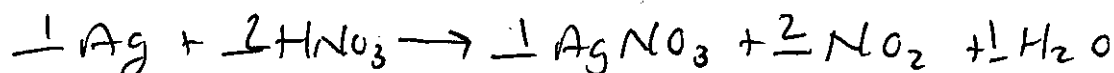
Iron (II) chloride reacting with chlorine to form iron (III) chloride.



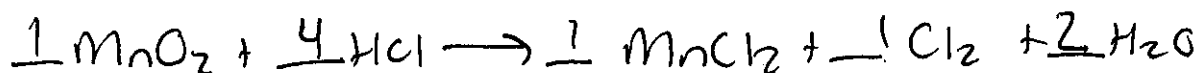
Ethanol ( $\text{C}_2\text{H}_5\text{OH}$ ) burning in air ( $\text{O}_2$ ) to form carbon dioxide and water.



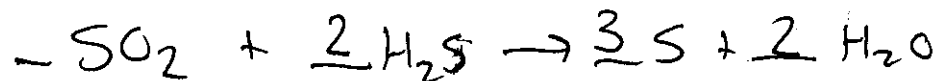
Silver reacting with nitric acid ( $\text{HNO}_3$ ) to form silver nitrate and nitrogen dioxide and water.



Manganese (IV) oxide reacting with hydrochloric acid ( $\text{HCl}$ ) to form manganese (II) chloride and chlorine and water.



Sulphur dioxide reacting with hydrogen sulfide ( $\text{H}_2\text{S}$ ) to form sulphur and water.



Ammonia ( $\text{NH}_3$ ) reacting with oxygen to form nitrogen monoxide and water.

